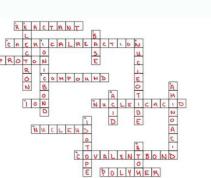


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**Chapter 3 – Chemistry of Life Test Review**

1. Review your notes, labs, and vocabulary for Chapter 2.
 2. Review your basic chemistry worksheets.
 3. Review the parameters and properties lab.
 4. Make sure you have review of angles and the Starkland lab!
- Sugar = Keweenaw's solution**
Lipid = Boston EJ
Amphiphile = Langford's solution
Polymer = Maruya
5. What are the different parts of an atom? what are their charges?
Proton = positive
Neutron = neutral
Electron = negative
 6. What does the symbol of the atom mean?
Protons & electrons
 Below electrons and protons. Write out the neutron and protons for the following equation:

Mg + Cl → MgCl

Magnesium = positive

7. Using a periodic table how many protons, neutrons, and electrons does Lithium have? What is the atomic number? What is the mass number?
Atomic number = number of protons
Mass number = sum of the protons and neutrons in the nucleus
8. Define atoms, elements, and molecules.
Atom = smallest repeatable entity that has a chemical identity
Element = substance made from one type of atom
Molecule = group of atoms bonded together
9. Give the electron form of H, big enough for it! BaCl₂
Ag = element
NaCl = molecule
10. What does nitrogen fixation allow plants to take up nitrogen in the soil?
Nitrate bond
11. What is carbon? Give example of? Why?
Isotope = different form of an element
12. Charge on acids and bases formation they will make? Give examples
 $\text{H} \rightarrow \text{2 acidic, 2 neutral, 2+10 basic, 1+3 strong acids, 11-14 strong bases}$

Memory Test Study Guide**Introduction Materials and Cells**

Materials consist of physically separate or chemically bonded structures of various sizes and shapes.

- Carbon monocovalent bonds, which allow great flexibility in forming complex linear and cyclic structures.
- Nonmetallic elements covalent bonding that are important for maintaining the structure of the bonds between and between molecules.
- Separation bonds include:
 - a) ionic bonding
 - b) Van der Waals
 - c) Hydrogen bonding
 - d) Dipole-dipole interactions
- Biomolecules are classified by combinations of different covalent bonds.

Biomolecules**Type of molecule**

- | | |
|--------------------|-----------------|
| - Polymers | - Lipoproteins |
| - Carbohydrates | - Proteins |
| - Lipids | - Lipids |
| - glycosidic bonds | - peptide bonds |

Building blocks**Subunit**

- | | |
|-------------------|----------------|
| - nucleotides | - Lipoproteins |
| - amino acids | - Proteins |
| - monosaccharides | - Lipids |
| - fatty acids | - Lipids |
| - glucose | - Lipids |

Carbohydrates heteropolymers, with a central chain and branch off containing a variety of monosaccharides. Disaccharides contain two monosaccharides at their linkage positions, which are linked together via hydrogen bonding.

Hydrophilicity

The hydrophilicity of some carbons are high (polar) and some are low due to molecular orientation, rotation.

- (a) there is a polar molecule. The partial negative charge on the oxygen (O) and the partial positive charge on the hydrogen (H) of water molecule.
- (b) hydrophilicity is caused by ionization because without ionization, these molecules form clusters from opposite charges, making themophobic (water) i.e. hydrophobic (water) and nonpolar molecules.

Please be an environmental stewardship.

Section 2-1 The Nature of Matter (pages 25-39)
 This section identifies the three particles that make up atoms. It also explains how atoms of the same element can have a different number of neutrons and describes the two main types of chemical bonds.

Answers (page 25)
 1. The basic unit of matter is called an atom.
 2. At the center of an atom is made up of protons and neutrons.

3. Complete the table about subatomic particles.
SUBATOMIC PARTICLES

Particle Charge Location in Atom
 Proton + Positive Nucleus
 Neutron 0 Neutral Nucleus
 Electron - Negative Surrounding nucleus

4. Why are atoms neutral despite having charged particles?

Atoms have equal numbers of electrons and protons, and these subatomic particles have equal but opposite charges.

Elements and Isotopes (page 30)

5. What is a chemical element?

A chemical element is a pure substance that consists entirely of one type of atom.

6. What does an element's atomic number represent?

It represents the number of protons in an atom of the element.

7. Atoms of the same element that differ in the number of neutrons in their nuclei are known as isotopes.

8. How are isotopes identified? Isotopes are identified by their mass numbers.

9. Why do all isotopes of an element have the same chemical properties? They have the same chemical properties because they have the same number of electrons.

Chemical Compounds (page 37)

10. What is a chemical compound?

A chemical compound is a substance formed by the chemical combination of two or more elements in definite proportions.

11. What does the formula for table salt indicate about that compound? The formula for table salt, NaCl, indicates that the elements that make table salt form a sodium and chlorine 1:1 ratio.

Chemical Bonds (page 38-39)

Date _____ Period _____ Name _____
 Chapter 2 Test Review Homework # 2 Answers

Match the mathematical expression with the name of the law.

1. Boyle's Law a. $P_1V_1 = P_2V_2$

2. Charles' Law b. $P_1V_1/T_1 = P_2V_2/T_2$

3. Gay-Lussac's Law c. $P_1 + P_2 = P_{\text{total}}$

4. Avogadro's Law d. $P_1V_1 = nRT_1$

5. Raoult's Law e. $\frac{P_A}{P_{\text{total}}} = \frac{n_A}{n_{\text{total}}}$

6. Solve the following problems. Show all work.

1. A sample of gas occupies 0.05 L at 150 K and 20K Pa. Find its volume at 273K.

2. What happens to pressure of the gas in a vessel whose temperature is increased from 20°C to 150°C?

3. If you bring a sealed bag of potato chips with you on airplane, the bag may burst if the cabin pressure drops too low. Explain why this is likely to happen during the flight. Why do most airlines prohibit passengers from bringing bags of chips on airplanes?

4. You have a 15 L container holding 10 g of helium at 1.0 atm. Find the new volume of the container if the atmospheric pressure is 0.75 atm. (Assume no change in temperature.)

5. How many grams of hydrogen (molar mass = 2.02 g/mol) are needed to produce 400 mL of H₂ gas at 1.00 atm and 25°C?

6. If you have a 400 mL container holding 10 g of helium at 1.0 atm. Find the new volume of the container if the atmospheric pressure is 0.75 atm. (Assume no change in temperature.)

7. If you have a 15 L container holding 10 g of helium at 1.0 atm. Find the new volume of the container if the atmospheric pressure is 0.75 atm. (Assume no change in temperature.)

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