
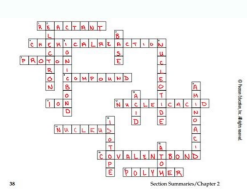


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**Chapter 2 - Chemistry of Life Test Review**

1. Review your notes, lab, and vocabulary for Chapter 2
2. Review your basic chemistry worksheet!
3. Review the macromolecule and organic lab!
4. Make sure you know names of layers of the bilbiklab lab!  
 Sugar - **Karbohydrat's valditiem**  
 Lipid - **lipidat's valditiem**  
 Protein - **proteinat's valditiem**
5. What are the different parts of an atom, what are they charged?  
 Proton - **positive**  
 Neutron - **neutral**  
 Electron - **negative**
6. What does the number of the atom mean?  
**Proton's electrons**
7. Write reactants and products. Where are the reactants and products in the following equation?  

$$2H_2 + O_2 \rightarrow 2H_2O$$
 Reactants - **prothins**
8. Using a periodic table how many protons, neutrons, and electrons does lithium have? What is the atomic number? What is the mass number?  
 Atomic number - **number of protons**  
 Mass number - **sum of the protons and neutrons in the nucleus**
9. Define atoms, elements, and molecules.  
 Atoms - **smallest type of matter that forms a chemical element**  
 Elements - **substance only made from one type of atom**  
 Molecules - **group of atoms bonded together**
10. How do atoms form a molecule? Give an example.  
 Ag - **element**  
 NaCl - **molecule**
11. What does an ionic bond form when it loses or gains an electron?  
**ionic bond**
12. What is Calcium (Ca) an example of? Why?  
**isotope, differs and amount of neutrons**
13. What are acids and bases? Give an example of each? Give examples.  
 H - 2 acid, 2 neutral, 7 - 10 basic, 1 - 2 strong acids, 11 - 14 strong bases

**Macromolecular Chemistry**

**Introduction: Molecules and Cells**

Macromolecular chemistry is the study of a molecule's structure and properties.

It is the study of the structure and properties of molecules.

- Carbon atoms combine to form **organic molecules**, with a strong preference for being complexed and cyclic structures.
- Macromolecular bonds are the **intermolecular forces** that are important for maintaining the structure of the complex molecules and their biological functions.
- Macromolecular bonds include:
  - i. Hydrogen bonds
  - ii. Hydrophobic
  - iii. Ionic interactions
  - iv. Hydrophobic interactions
- Molecules with a strong preference for being complexed and cyclic structures.

**Macromolecular Chemistry**

- Protein

- Lipids

- Carbohydrates & Polymers

- Nucleic acids

- Amino acids

- Macromolecules

- Lipid acids

- Glucose

**Macromolecular Chemistry**

- Nucleic acids

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- Carbohydrates & Polymers

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- Macromolecules

- Lipid acids

- Glucose

Macromolecular chemistry is the study of a molecule's structure and properties. It is the study of the structure and properties of molecules.

**Water and Solids**

The unique properties of water (such as high boiling point and melting point) are due to its molecular structure, **hydrogen bonds**.

- i. Water is a polar molecule. The partial negative charge on the oxygen (O) and the partial positive charge on the two hydrogen (H) atoms attract.
- ii. Hydrogen bonds are formed by interactions between different H<sub>2</sub>O molecules. These molecules are held together by weakly changing, weakly hydrogen-bonded network. It is the hydrogen bonding which stabilizes structure.

Water is an essential nutrient for many organisms.



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